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Assessing Peroxygenase-Mediated Oxidations in the Presence of High Concentrations of Water-Miscible Co-Solvents

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The use of water-miscible organic co-solvents in biocatalysis is a simple procedure for obtaining higher enzymatic activities toward hydrophobic substrates. However, effects on activity and stability have to be carefully evaluated, also with regard to the type and concentration of the respective co-solvent. In this contribution, we investigated and evaluated the effect of some common water-miscible co-solvents on the biocatalytic performance of the recombinant unspecific peroxygenase rAaeUPO from *Agroclybe aegerita*. rAaeUPO showed promising activities in the presence of high concentrations of the best co-solvent acetonitrile, which enabled to use higher substrate concentrations (≥ 100 mM). Employing high acetonitrile concentrations for UPO-mediated oxidation of ethylbenzene to (*R*)-1-phenylethanol was demonstrated under preparative scale conditions and led to product accumulation rates of 31 mM h^{-1} .

Keywords: peroxygenase, oxyfunctionalization, co-solvent, hydroxylation, epoxidation, medium engineering

INTRODUCTION

Biocatalysis bears an enormous potential for the chemical industry to become more resource-saving and environmentally more benign (Ni et al., 2014; Sheldon and Woodley, 2018). However, to tap its full potential, a range of issues still have to be addressed. The hydrophobicity of most reagents of interest represents one of these challenges. As aqueous buffer systems still represent the reaction media of choice for biocatalysis, frequently, only low reagent concentrations can be realized. In fact, today, the majority of biocatalytic reactions reported by academic research groups deal with reagent concentrations below 100 mM (Holtmann and Hollmann, 2022). Typically, this corresponds to maximal product loadings of less than 15 g L^{-1} and wastewater amounts for more than $60 \text{ kg kg}_{\text{product}}^{-1}$. Obviously, this situation is far away from being economically and environmentally sustainable!

Among the various approaches to increase the reagent concentrations, water-miscible organic solvents are the most widespread strategies (van Schie et al., 2021). The presence of such co-solvents, however, can exceed unpredictable effects on the biocatalyst. For various enzymes, co-solvents have been reported to (negatively) influence their activity and stability (Klibanov, 2001; Serdakowski and Dordick, 2008; Doukyu and Ogino, 2010; Stepankova et al., 2013; Dutta Banik et al., 2016).

Fungal peroxygenases are heme-thiolate enzymes catalyzing a broad range of preparatively interesting conversions such as regio-, chemo- and enantioselective oxyfunctionalization reactions (Hobisch et al., 2020a). As extracellular enzymes, peroxygenases may be expected to be robust also under harsh reaction conditions such as the presence of high concentrations of organic solvents.

Thus, the use of peroxygenases in two-liquid phase systems (Churakova et al., 2011; Tonin et al., 2021) or neat reaction media (Poraj-Kobielska et al., 2015; Fernández-Fueyo et al., 2016; Rauch et al., 2019; Hobisch et al., 2020b) has already been demonstrated, whereas the application in one-phase systems, in particular with high amounts of simple water-miscible solvents, still needs to be explored. In the current contribution we investigate the effect of a range of water-miscible organic co-solvents on fungal peroxygenases, using the recombinant evolved peroxygenase from *Agrocybe aegerita* (rAaeUPO) as a model enzyme.

MATERIALS AND METHODS

Preparation of Recombinant Unspecific Peroxygenase From *Agrocybe aegerita*

The expression-engineered variant rAaeUPO (PaDa-I mutant) was typically used as concentrated supernatant or purified enzyme derived from a 2,500 L pilot-scale cultivation of recombinant *Pichia pastoris* X-33 expressing rAaeUPO (Molina-Espeja et al., 2014; Tonin et al., 2021). Purification was similarly performed according to previous studies (Molina-Espeja et al., 2014). rAaeUPO concentrations were determined in CO-difference spectra (see below) or *via* their absorption at 420 nm using the extinction coefficient ($\epsilon_{420} = 115 \text{ mM}^{-1} \text{ cm}^{-1}$) (Fernández-Fueyo et al., 2016).

CO-Difference Spectra

The amount of rAaeUPO in concentrated supernatants was determined using CO-difference spectra. 950 μL of protein sample, diluted in 100 mM KPi-buffer if necessary, were filled into plastic cuvettes and placed in a photometer. After recording a blank (base subtraction), the sample was exposed with CO for a few seconds. Next, 50 μL of a 1 M sodium dithionite stock solution was added, and a difference spectrum between 400 and 500 nm was recorded. The measurements were continued until a constant absorption maximum was obtained. All measurements were performed in duplicates. The concentration of rAaeUPO was calculated based on the absorption values at 445 and 490 nm using the molar extinction coefficient of $\epsilon_{445} = 107 \text{ mM}^{-1} \text{ cm}^{-1}$ (Tieves et al., 2019a).

ABTS Assay

The peroxidase activity of rAaeUPO was determined in a continuous photometric assay against the substrate 2,2'-Azino-bis(3-ethylbenzothiazoline-6-sulfonic acid (ABTS) at 420 nm ($\epsilon_{420} = 36.0 \text{ mM}^{-1} \text{ cm}^{-1}$). AaeUPO (50 μL) in appropriate dilution (typically 50- or 100-fold) was mixed with 0.5 mM ABTS in 100 mM sodium citrate buffer, pH 4.4 in 2 mL plastic cuvettes. After short incubation at 25°C, measurements were started by the addition of 2 mM H_2O_2 . The increase of absorption at 420 nm caused by ABTS formation was tracked for 90 s at 25°C in a Cary 60 UV-VIS spectrophotometer (Agilent). The initial slope ($\Delta A_{420}/\text{min}$) between 20–80 s was linear and used to calculate the volumetric activity [U/mL]. 1 U is

defined as the amount of enzyme which is needed to convert 1 μmol substrate in 1 min under assay conditions. All measurements were performed in triplicates.

Purpald Assay

The activity of rAaeUPO against the primary alcohols methanol and ethanol was examined using the purpald assay (Dickinson and Jacobsen, 1970; Hopps, 2000). Reaction mixtures contained 1 μM rAaeUPO (concentrated supernatant or purified enzyme), 20 vol% co-solvent, and H_2O_2 , which was added with a dosing rate of 4 mM h^{-1} . The 200 mM H_2O_2 -stock solution was freshly prepared from commercially available H_2O_2 (50 wt% solution, Merck). For analysis, 5–50 μL sample was diluted in 100 mM potassium phosphate buffer, pH 7.5 to a final volume of 200 μL in a 96-well plate. Afterward, 50 μL purpald (0.16 M in 2 M NaOH) was added, and samples were properly mixed. The samples were incubated at room temperature for at least 1 h, allowing the formation of a purple dye in the presence of aldehydes and oxygen.

In Vitro Conversions

rAaeUPO-mediated oxidation reactions were typically performed in 0.5 mL 100 mM potassium phosphate buffer, pH 6.0 in 2 mL glass vials. Reaction mixtures contained 1 μM rAaeUPO (concentrated supernatant), 10 mM ethylbenzene or *cis*- β -methyl styrene (in 10–97 vol% of acetonitrile), and H_2O_2 , which was added with a dosing rate of 4 mM h^{-1} . The 200 mM H_2O_2 -stock solution was freshly prepared from commercially available H_2O_2 (50 wt% solution, Merck). The samples were incubated at 25°C and 600 min^{-1} for 0–3 h in an Eppendorf shaker.

In order to track the substrate oxidation with automated H_2O_2 addition, reactions in a 10 mL scale were prepared in glass vessels (Supplementary Figure S5). The glass devices were placed in an Eppendorf shaker equipped with a 50 mL falcon attachment, which was fixed in such a way that proper shaking and correct placement of the glass vials were enabled as it would be the case for 50 mL falcons. Spring pumps containing a stock solution of 800 mM H_2O_2 (freshly prepared from a 50 wt% H_2O_2) were connected and automatically pumped 100 μL of H_2O_2 every hour (dosing rate: 8 $\text{mM h}^{-1} \text{H}_2\text{O}_2$).

Ethyl acetate (500 μL) containing 5 mM of the internal standard *n*-dodecane was used for extraction. After addition, samples were vortexed and centrifuged at 13,000 $\times g$ for 3 min at room temperature. The organic phase was transferred to a plastic vial with inlet and used for GC analysis.

Kinetic Studies

The kinetic parameters K_m and k_{cat} for the rAaeUPO substrates ethylbenzene and 1-phenylethanol were determined in the presence of different concentrations of the co-solvent acetonitrile. Reaction mixtures (8 mL) in 10 mM KPi buffer (pH 7) contained 8.4 nM rAaeUPO (concentrated supernatant), 1.7 mM H_2O_2 (freshly diluted from a 50 wt% H_2O_2 stock solution), and the substrates ethylbenzene or 1-phenylethanol at concentrations ranging from 0.03–2 mM. Acetonitrile concentrations varied from 0–20 vol%. For

reactions without acetonitrile, the maximum concentration of ethylbenzene was set at the solubility limit of 1.2 mM (150 mg/L, IPCS Inchem.). Samples were incubated at room temperature for 15 s, and then the reaction was stopped by adding 1 mL 50 vol% trifluoroacetic acid. Afterward, the samples were extracted with ethyl acetate containing 5 mM of the internal standard 1-octanol, dried over magnesium sulfate, and analyzed by GC. K_m and v_{max} values were obtained by fitting the reaction rate (μM product per min) vs. the substrate concentration according to the equation of Michaelis–Menten by non-linear regression using Prism GraphPad, version 5 (Supplementary Figures S9–S11). The turnover number k_{cat} was calculated by dividing v_{max} by the total UPO concentration.

Application of rAaeUPO Under Preparative Scale Conditions

The rAaeUPO-mediated hydroxylation of ethylbenzene on a 100 mL scale was performed in a SYSTAG jacketed lab reactor (250 mL operational volume) at 25°C and 300 rpm shaking speed. Reaction solution (100 mL) contained 100 mM KPi buffer (pH 6), 50 vol% acetonitrile, 10 μM rAaeUPO (concentrated supernatant), and 300 mM ethylbenzene. H_2O_2 was added continuously from a 4.5 M stock solution with a H_2O_2 -dosing rate of 50 mM h^{-1} and was freshly prepared from a commercial 50 wt% H_2O_2 solution (Merck) prior to the experiment. At different time points, samples from the aqueous phase were withdrawn, extracted with ethyl acetate containing 5 mM *n*-dodecane (IS), and analyzed *via* chiral GC.

Product Analysis

For GC analysis, a Shimadzu GC-14A/FID or Shimadzu GC-2010 plus/FID system was used. Achiral GC-analysis was performed on an Agilent CP-Sil 5 CB column (50 m \times 0.53 mm \times 1.0 μm) with N_2 as carrier gas. Kinetic studies were analyzed on a CP-Wax 52 CB column (25 m \times 0.25 mm \times 1.2 μm) using N_2 as carrier gas. Chiral GC-analysis for *cis*- β -methyl styrene and products was performed according to Rauch et al. on a Lipodex E column (Macherney Nagel) using helium as carrier gas (Rauch et al., 2019). Ethylbenzene and oxidized products thereof were analyzed on a CP-Chirasil-Dex CB (Agilent) (25 m \times 0.32 mm \times 0.25 μm) with helium as carrier gas (Tieves et al., 2019b). Substrate and product concentrations were calculated based on calibration curves. Because an authentic standard for (1*R*,2*S*)-*cis*-methyl styrene oxide was not available, the response factor between calibration curves of styrene and styrene epoxide was used to calculate the concentration of (1*R*,2*S*)-*cis*-methyl styrene oxide relative to *cis*- β -methyl styrene.

Temperature programs are described in key points as follows:

GC program achiral analysis (Split 10)

100°C hold 3 min and 20°C/min to 325°C hold 3 min; retention times: 3.18 min *cis*- β -methyl styrene and 4.5 min *cis*-methyl styrene oxide; potential minor side product peaks (not in negative control): 4.76, 6.08, and 6.42 min; retention times: 5.1 min ethylbenzene, 7.5 min 1-phenylethanol, and 7.7 min acetophenone; potential side product peaks (not in negative control): 5.45 min and IS *n*-dodecane 8.8 min.

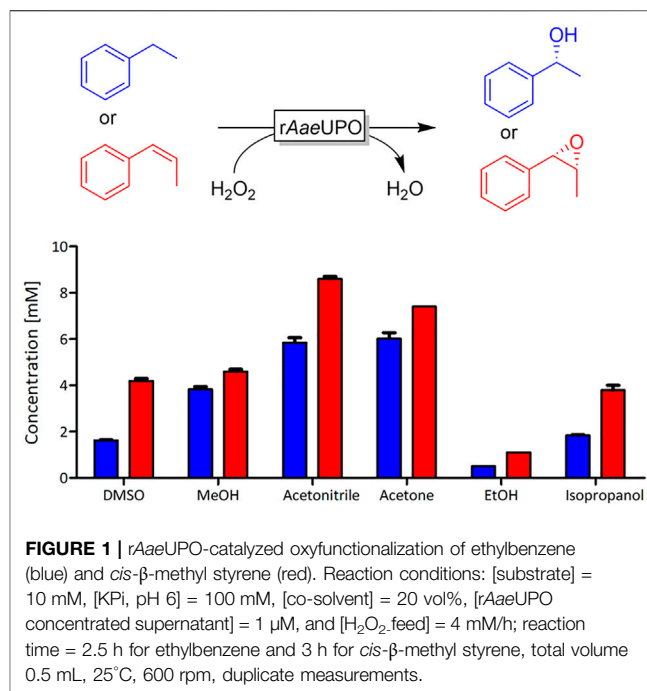


FIGURE 1 | rAaeUPO-catalyzed oxyfunctionalization of ethylbenzene (blue) and *cis*- β -methyl styrene (red). Reaction conditions: [substrate] = 10 mM, [KPi, pH 6] = 100 mM, [co-solvent] = 20 vol%, [rAaeUPO concentrated supernatant] = 1 μM , and [H_2O_2 .feed] = 4 mM/h; reaction time = 2.5 h for ethylbenzene and 3 h for *cis*- β -methyl styrene, total volume 0.5 mL, 25°C, 600 rpm, duplicate measurements.

GC program “kinetics” for ethylbenzene or 1-phenylethanol (Split 150)

135°C hold 4 min, 20°C/min to 195°C hold 5 min, and 30°C/min to 250°C hold 1 min; retention times: 3.3 min ethylbenzene, 9.4 min acetophenone, 11.5 min 1-phenylethanol, and IS 1-octanol 7.5 min.

GC program chiral analysis for *cis*- β -methyl styrene (Split 100)

100°C hold 15 min and 20°C/min to 220°C hold 1 min; retention times: 5.44 min *cis*- β -methyl styrene and 10.77 min (1*R*,2*S*)-*cis*-methyl styrene oxide.

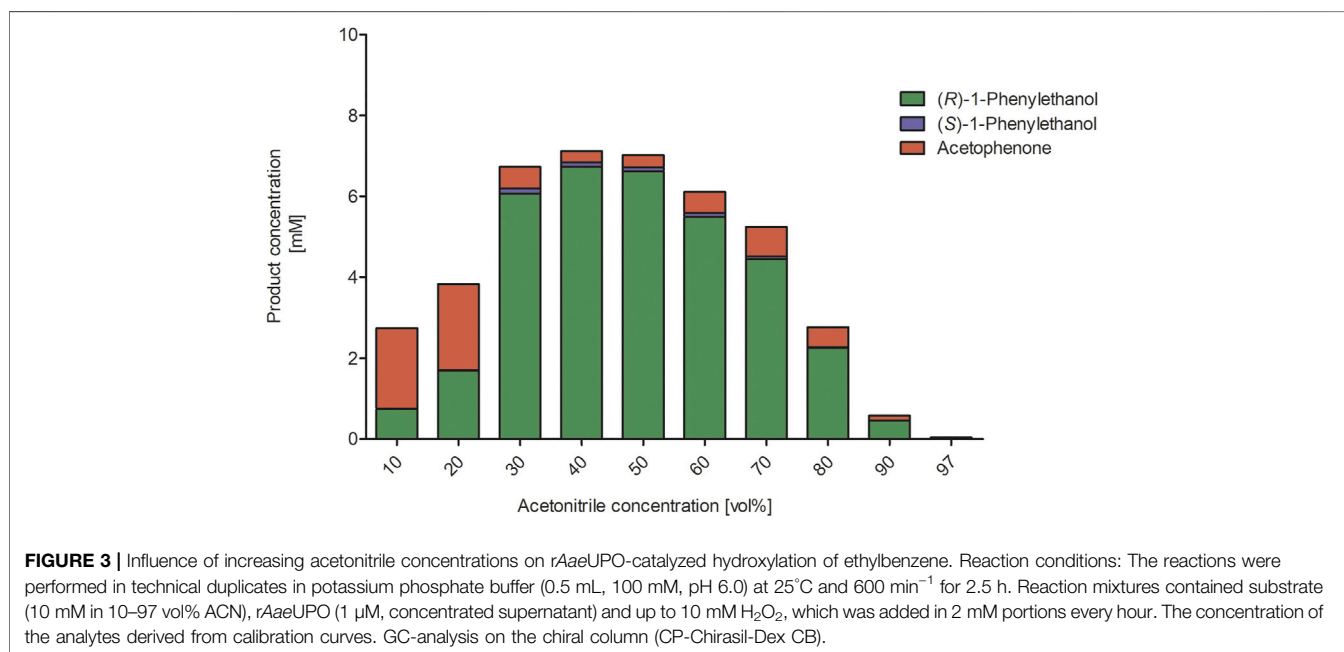
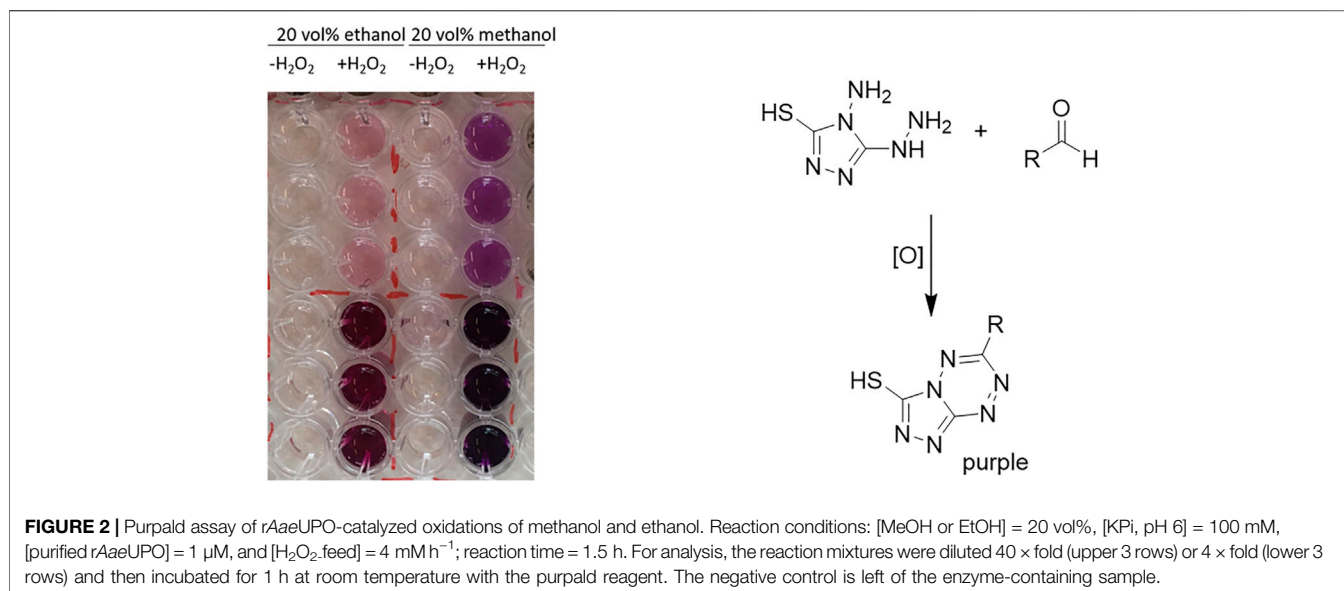
GC program chiral analysis for ethylbenzene (Split 50)

120°C hold 2.6 min, 15°C/min to 135°C hold 3.3 min, and 25°C/min to 225°C hold 1.0 min; retention times: 2.47 min ethylbenzene, 6.38 min (*R*)-1-phenylethanol, 6.65 min (*S*)-1-phenylethanol, 4.30 min acetophenone, and IS *n*-dodecane 4.58 min.

RESULTS AND DISCUSSION

Effect of Water-Miscible Co-solvents on the Activity and Selectivity of rAaeUPO

To investigate the effect of water-soluble organic co-solvents on the activity of rAaeUPO, we evaluated a selection of commonly used solvents ranging from dimethyl sulfoxide (DMSO) ($\log P = -1.35$) to isopropanol ($\log P = 0.05$) at a substrate concentration of 10 mM. As model reactions, we chose the hydroxylation of ethylbenzene and the epoxidation of *cis*- β -methyl styrene (Figure 1) in the presence of 20 vol% of the respective



solvents. It is worth mentioning that the enantioselectivity of both reactions was not significantly impaired by the co-solvent (**Supplementary Figures S1, S2**).

Both model reactions were influenced by the presence of the solvents in a similar fashion. There was no obvious correlation between the polarities of the solvent (as expressed by their logP value). Specially for the alcohol-based solvents methanol (MeOH) and ethanol (EtOH), we suspected that possibly alcohol oxidation may compete with the model reaction investigated. Note that the molar surplus of alcohol solvent over the starting material was approximately 250-fold. In fact, using the purpald assay (Dickinson and Jacobsen, 1970; Hopps,

2000), we could qualitatively confirm the formation of formaldehyde and acetaldehyde from the corresponding alcohol oxidations catalyzed by rAaeUPO (**Figure 2**). We assume that also isopropanol at least to some extent can be converted by rAaeUPO but an analytical validation is yet pending. Also in case of DMSO, which is an inhibitor of peroxygenase activity, GC/MS analysis confirmed rAaeUPO-conversion of the solvent to the sulfone (**Supplementary Figure S6**). For acetone and acetonitrile, we did not observe any indication for co-solvent oxidation in our experiments and nearly full conversion ($\geq 90\%$) of 10 mM substrate was achieved under the experimental conditions chosen.

TABLE 1 | Effect of increasing acetonitrile (ACN) concentrations on the kinetic parameters of rAaeUPO toward the substrates ethylbenzene or 1-phenylethanol. Reaction conditions: [ethylbenzene or 1-phenylethanol] = 0.3–2 mM, [KPI, pH 7] = 10 mM, [acetonitrile] = 0–20 vol%, [rAaeUPO concentrated supernatant] = 8.4 nM, [H₂O₂-concentration] = 1.7 mM; reaction time = 15 s, total volume 8 mL, room temperature, and triplicate measurements.

Substrate	ACN-concentration [vol%]	K _m [mM]	k _{cat} [s ⁻¹]
Ethylbenzene	0	0.7 ± 0.1	415 ± 32
Ethylbenzene	10	0.8 ± 0.1	470 ± 27
Ethylbenzene	20	0.6 ± 0.1	409 ± 24
1-Phenylethanol	0	1.4 ± 0.2	170 ± 10
1-Phenylethanol	10	2.7 ± 0.4	455 ± 41
1-Phenylethanol	20	5.8 ± 1.4	643 ± 106

Activity and Stability of rAaeUPO in the Presence of High Acetonitrile Concentrations

We next investigated the influence of the acetonitrile concentration on the rAaeUPO-catalyzed hydroxylation reaction (Figure 3). Increasing the acetonitrile concentration up to approximately 50 vol% had a beneficial effect on product formation. Notably, even in the presence of very high acetonitrile concentrations of 90 vol% significant activity was detected. Possibly, the increasing product accumulation up to 50 vol% can be attributed to the increasing solubility of the reactants in the reaction medium thereby also minimizing the undesired evaporation. With increasing acetonitrile concentrations, also a larger amount of the alcohol product (1-phenylethanol) compared to the double oxidation product acetophenone was formed. We decided to investigate this effect by further determining the kinetic parameters K_m and k_{cat} for the substrates ethylbenzene and 1-phenylethanol at different acetonitrile concentrations (Table 1). It was found that increasing acetonitrile concentrations did not significantly affect the kinetic parameters for the substrate ethylbenzene. In contrary, higher K_m and higher k_{cat} values for the substrate 1-phenylethanol were observed when the acetonitrile amount was increased. Currently, we are lacking a full understanding of this behavior as the better solubility of the reactants as well as the solvent presence in the active site might contribute to these changes in kinetic parameters and the product concentrations, respectively. Thus, further investigations in this regard need to be performed. Nevertheless, the use of increased acetonitrile concentrations seems attractive in order to avoid over oxidation to the ketone and obtain an accumulation of the desired hydroxylation product.

Increasing the co-solvent concentration beyond 50 vol% most likely has a negative impact on the structural integrity of the biocatalyst explaining the decreasing product accumulation (Figure 3). Similar observations were also made for the epoxidation of *cis*-β-methyl styrene, indicating that these effects are independent on the tested substrate (Supplementary Figure S7). To further investigate the influence of acetonitrile on the stability of rAaeUPO, we performed a range of pre-incubation experiments exposing the

enzyme to varying concentrations of acetonitrile and determining at intervals the residual activity (Figure 4). While the enzyme, incubated in buffer without any acetonitrile, showed no significant decrease in catalytic activity over several days, a sharp drop was observed in the presence of 50 and 80 vol% acetonitrile, resulting in approximately 40 and 1% of residual enzyme activity after 1 h. In the presence of 50 vol% acetonitrile, rAaeUPO activity did not decrease further for several days. The dramatic decrease in catalytic activity was accompanied by a green precipitate evolving in the incubation mixture. We assume that the dramatic activity losses in the first few hours of incubation can be attributed to precipitation of the biocatalyst. Activity measurements of the recovered precipitate from 50 vol% acetonitrile resuspended in buffer indicated that the precipitated UPO was still active (Supplementary Figure S12). Thus, the loss of activity in 50 vol% acetonitrile is most likely not due to enzyme inactivation but rather to a decreased solubility of the enzyme in the acetonitrile/buffer mixture. Consequently, future studies might provide exciting possibilities to use acetonitrile for UPO precipitation to easily obtain active enzyme immobilisates and compare them to previous immobilization procedures (Rauch et al., 2019; Hobisch et al., 2020b).

Complementary to the stability measurements, we followed the rAaeUPO-catalyzed oxidation of the model substrate ethylbenzene over time in the presence of high concentrations of acetonitrile (Figure 5). We observed a different effect of the tested acetonitrile concentration of 50 and 80 vol%. The product concentration increased linearly within 7 h in the presence of 50 vol% acetonitrile, whereas the reaction stopped earlier in the presence of 80 vol% acetonitrile. This resulted in a maximum target product concentration of 57 mM (*R*)-1-phenylethanol in 80 vol% acetonitrile. Looking at the stability measurements (Figure 4), the lower product formation in the presence of 80 vol% acetonitrile can be attributed to insufficient enzyme activity and stability in the presence of co-solvent. In this regard, it is still astonishing that over 50 mM (*R*)-1-phenylethanol are formed though the enzyme activity was less than 1% compared to its initial value according to our stability measurements. Using 50 vol% acetonitrile for the UPO-mediated oxidation of *cis*-β-methyl styrene resulted in similar trends, indicating that these conditions can also be applied for other UPO substrates (Supplementary Figure S8).

Preparative-Scale Application and E-Factor Analysis

Encouraged by these promising results, we decided to demonstrate the preparative-scale application using an ethylbenzene starting concentration of 300 mM (Figure 6). The reaction produced smoothly accumulating enantiomerically pure (*R*)-1-phenylethanol until nearly all starting material was consumed after 6 h. Only then, over oxidation to acetophenone occurred. Within 6 h, 188 mM of enantiomerically pure (*R*)-1-phenylethanol has been formed corresponding to a turnover number (TN = mol_{Product} × mol_{rAaeUPO}⁻¹) of 18.800 and a turnover frequency (TOF) of 0.9 s⁻¹ (over 6 h). These turnover numbers were lower

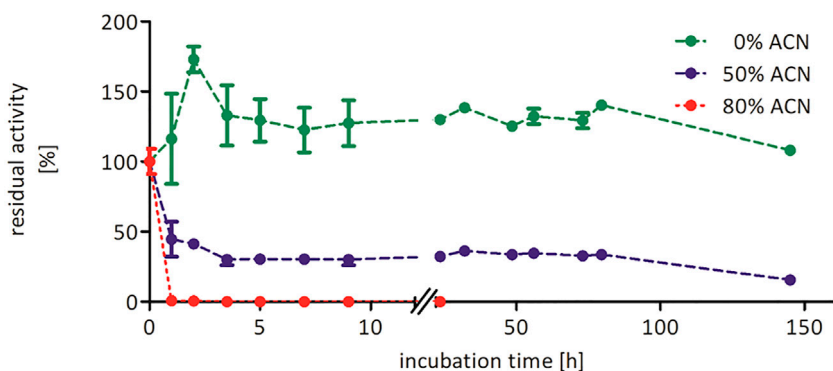


FIGURE 4 | Stability of rAaeUPO (concentrated supernatant) in the presence of 0 vol% (green), 50 vol% (blue), or 80 vol% (red) acetonitrile. Enzyme (1 μM) was incubated in 100 mM KPi, pH 6.0, and indicated acetonitrile concentrations at 25°C and 600 rpm in glass vials with 10 mL starting volume. At indicated time points, a 200 μL aliquot was transferred to an Eppendorf tube, and enzyme activity was evaluated via ABTS assay. Data are normalized to the 0 h sample, which was taken immediately after mixing ($t_0 = 27.3$ U/mL).

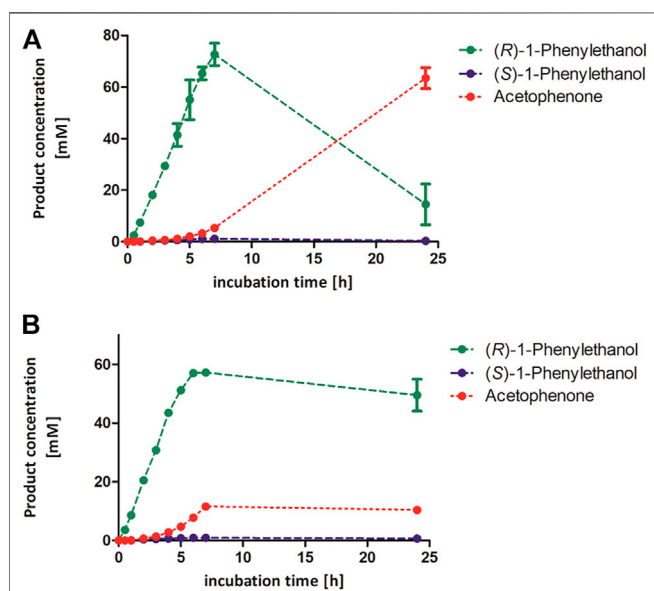


FIGURE 5 | Product concentrations during conversion of 100 mM ethylbenzene by rAaeUPO applying 50 vol% (A) or 80 vol% (B) acetonitrile. Reaction conditions: 100 mM potassium phosphate buffer, pH 6.0, 1 μM rAaeUPO (concentrated supernatant), and H_2O_2 -dosing rate: 8 mM h^{-1} . The concentration of both 1-phenylethanol enantiomers and acetophenone were calculated based on a calibration curve recorded earlier. Duplicates. GC-analysis was performed on the chiral column (CP-Chirasil-Dex CB).

compared to previous results from our group, where for the same substrate and enzyme in a two-liquid phase system TNs of 340.000 were achieved (Tonin et al., 2021). However, it should be pointed out that the achieved concentration of 120 mM (*R*)-1-phenylethanol was obtained after 7 days of incubation resulting in a low product accumulation rate of 0.7 mM h^{-1} . The initial (*R*)-1-phenylethanol accumulation rate in our experiments was more than 31 mM h^{-1} (3.8 $\text{g}_{\text{product}} \text{h}^{-1} \text{L}^{-1}$) and thus represent to the best of our knowledge one of the highest productivities reported for a biocatalytic oxyfunctionalization reaction to date.

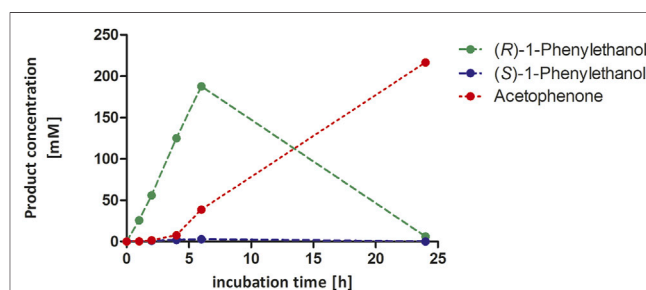


FIGURE 6 | Hydroxylation of ethylbenzene to (*R*)-1-phenylethanol in presence of 50 vol% acetonitrile in 100 mL scale. Reaction conditions: 300 mM substrate concentration, 100 mM potassium phosphate buffer, pH 6.0, 10 μM rAaeUPO (concentrated supernatant), and H_2O_2 -dosing rate: 50 mM h^{-1} . The reaction was stirred at 300 rpm at 25°C. The concentration of all analytes derived from calibration curves. GC-analysis was performed on the chiral column (CP-Chirasil-Dex CB).

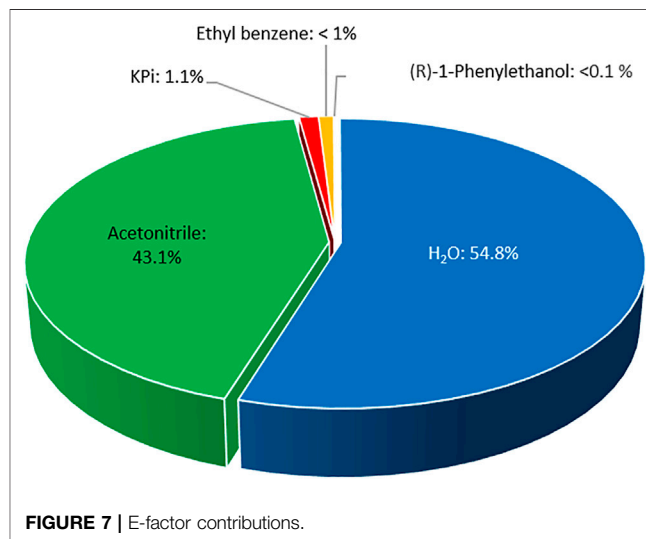


FIGURE 7 | E-factor contributions.

Overall 216 mM of acetophenone (double oxidation product) were obtained corresponding to 25.9 g L⁻¹ and 73% yield on the starting material used. The product to the catalyst ratio was 70 g_{product} g_{AaeUPO}⁻¹. As no starting material and only minor (*R*)-1-phenylethanol was detected after 24 h, we conclude that still evaporation of the reagents (especially the starting material) occurred. Nevertheless, these numbers bring peroxygenase-catalysis close to industrial requirements (Tufvesson et al., 2011).

From an environmental point of view, we believe that the application of acetonitrile represents an interesting alternative to pure aqueous media. A simple E-factor analysis (Sheldon, 2017; 2018) (Figure 7) revealed that in the experiment shown in Figure 6 approximately 35.2 g_{waste} g_{product}⁻¹ have been generated. 98% of this waste has been caused by the solvents (19.3 g_{water} g_{product}⁻¹ and 15.2 g_{acetonitrile} g_{product}⁻¹). The non-reacted starting material and phosphate buffer contributed roughly 1 g g_{product}⁻¹ each, whereas the contribution of the biocatalyst was almost negligible (<0.02 g_{AaeUPO} g_{product}⁻¹). Hence, the overall E-factor compared to similar reactions could be reduced more than 7-fold (Xu et al., 2022). In this regard, the direction for further improvements of the reaction's environmental footprint is also clear by further increasing the product concentration.

Obviously, a simple E-factor analysis does not replace a more detailed environmental impact analysis taking the history of the reagents used and energy contributions into account (Tieves et al., 2019a). Nevertheless, this analysis already demonstrates the impact of increasing product concentrations.

CONCLUSION

Water-miscible organic solvents represent a simple measure to increase the substrate- and product concentrations in biocatalysis reactions and thereby decrease the solvent wastes generated. In this contribution, we demonstrated that the peroxygenase from *Agroclybe aegerita* exhibits superb activity and stability in the presence of acetonitrile, enabling significantly increased reagent payloads in *AaeUPO*-catalyzed oxyfunctionalization reactions. This feature makes the use of this peroxygenase already remarkable as other heme-containing enzymes such as the peroxygenase *HspUPO* or the frequently investigated cytochrome P450 BM3 get already inactivated in presence of low concentrations of water-miscible co-solvents of typically less than 20 vol% (Seng Wong et al., 2004; Rotilio et al., 2021). Further improvements may be even

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expected using evolved variants of *rAaeUPO* with even higher solvent tolerance (Martin-Diaz et al., 2021).

The productivities and product concentrations achieved in preliminary experiments using high co-solvent concentrations point toward economically feasible and environmentally benign preparative applications.

DATA AVAILABILITY STATEMENT

The original contributions presented in the study are included in the article/Supplementary Material, further inquiries can be directed to the corresponding author.

AUTHOR CONTRIBUTIONS

TH planned, designed and carried out most experiments, analysed the data, and drafted the manuscript. AvT performed and evaluated kinetic experiments with different co-solvent concentrations. MA and FH acquired funding, gave advices in the research work, helped in drafting the manuscript, and revised the manuscript. All authors read and approved the final manuscript.

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SUPPLEMENTARY MATERIAL

The Supplementary Material for this article can be found online at: <https://www.frontiersin.org/articles/10.3389/fctls.2022.882992/full#supplementary-material>

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